

Concept of thermodynamic capacity

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Abstract. The thermodynamic capacity of a species (C_i) in a homogeneous phase is defined as $(\partial n_i / \partial \mu_i)_{p, T, n_j}$ where n_i is the total number of moles of i per unit quantity of the system irrespective of the actual system chemistry and μ_i is its chemical potential. Based on this definition, the thermodynamic capacity of oxygen in non-reactive and reactive gas mixtures and in binary and ternary liquid solutions has been computed. For reactive gas mixtures containing stable chemical species which do not undergo significant dissociation such as $\text{CO} + \text{CO}_2$, $\text{H}_2 + \text{H}_2\text{O}$ and $\text{H}_2 + \text{CO}_2$, the capacity curves show a maximum at equimolar ratio and a minimum at higher oxygen potentials. If one of the chemical species partly dissociates as in the case of H_2S in $\text{H}_2 + \text{H}_2\text{S}$ mixtures or SO_3 in $\text{SO}_2 + \text{SO}_3$ mixtures, capacity curves do not exhibit such maxima and minima, especially at high temperatures. It would be difficult to produce stable oxygen fugacities when the capacity has a low value, for example at compositions near the minimum. Oxygen capacities of non-ideal liquid solutions, Cu-O and Cu-O-Sn, and heterogeneous systems formed at saturation with the respective oxides are discussed.

Keywords. Thermodynamic capacity; ideal gas mixture; reactive gas mixture; condensed phase.

1. Introduction

While the concept of thermodynamic potential is well established and widely used in chemical literature, the corresponding capacity function has not been discussed in the textbooks on thermodynamics (Prigogine and Defay 1954; Swalin 1972; Kestin 1979; Lupis 1983). As such, chemical potentials are appropriate to determine the direction of transport of a species or of a reaction in a chemical system. However, being an intensity factor, chemical potentials do not provide any information on the amount of species transported or on the extent of reaction, even in the absence of kinetic restrictions. A suitably defined capacity factor can be useful in quantifying the amount of a species that can be drawn from a chemical reservoir at a specified chemical potential. Consider, for example, the synthesis of a non-stoichiometric oxide or a Magneli phase at high temperatures by equilibrating the solid with a flowing gas of defined oxygen potential. As oxygen is being added or removed from the oxide, the oxygen potential of the gas phase will usually change. If an $\text{Ar} + \text{O}_2$ mixture with $p_{\text{O}_2} < 10^{-4}$ atm is used for controlling the stoichiometry, the change in oxygen potential of the gas accompanying a finite oxygen exchange with the solid will be very large. In other words, the gas phase is unable to supply or remove any significant amount of oxygen from the solid at a defined potential. The oxygen capacity of the $\text{Ar} + \text{O}_2$ mixture may therefore be considered to be low at these low oxygen partial pressures. A gas mixture consisting of $\text{CO} + \text{CO}_2$ or $\text{H}_2 + \text{H}_2\text{O}$ may have higher capacity to supply or remove oxygen at low potentials. The purpose of

this communication is to define thermodynamic capacity and to illustrate its variation for some typical systems. A potential application of the concept in high temperature material synthesis and solid state device research is projected.

2. Definition of thermodynamic capacity

Just as the electrical potential increases as electrons are put into a condenser, the chemical potential of a species increases in a homogeneous medium with the addition of that species. While the electrical potential rises linearly with charge, the chemical potential increases logarithmically with concentration for ideal or dilute solutions. The thermodynamic capacity of a species (*i*) in a homogeneous phase may be defined by the expression,

$$C_i = (\partial n_i / \partial \mu_i)_{P, T, n_{j(j \neq i)}} \quad (1)$$

where n_i is the total moles of *i* in any state of chemical combination per unit amount of the phase and μ_i is the chemical potential of *i*. Thermodynamic capacity of species *i* may be visualized as the moles of *i* that is added or removed from a specified quantity of a homogeneous system when the chemical potential of *i* is given a unit increment or decrement.

It is interesting to explore the variation of thermodynamic capacity with composition for an ideal gas mixture (Ar + O₂), reactive gas mixtures usually used in chemical synthesis (CO + CO₂, H₂ + H₂O, SO₂ + SO₃, H₂ + CO₂ and H₂ + H₂S) and condensed phase solutions (Cu-O and Cu-O-Sn). For simple mixtures, analytical expressions may be derived relating capacity and composition. However, in systems containing several complex species, numerical procedures are more convenient. In the numerical technique, one starts with a preset composition and computes the chemical potential of species *i*. The chemical potential is then incremented by a small amount and the new composition is computed. From the change in composition, the amount of the species added is calculated. The capacity

$$C_i = \Delta n_i / \Delta \mu_i.$$

3. Capacity of an ideal gas mixture

An example of a simple gas mixture is Ar + O₂. In the ideal gas approximation, chemical potential of oxygen in the mixture is given by

$$\mu_{O_2} = \mu_{O_2}^{\circ} + RT \ln p_{O_2}. \quad (2)$$

Implicit differentiation gives

$$d\mu_{O_2} = (RT/p_{O_2}) dp_{O_2}. \quad (3)$$

The partial pressure of oxygen is defined by

$$p_{O_2} = n_{O_2} / (n_{O_2} + n_{Ar}). \quad (4)$$

This on differentiation gives

$$dp_{O_2} = (p_{Ar}/n) dn_{O_2} = (p_{Ar}RT/PV) dn_{O_2}, \quad (5)$$

where $n = n_{O_2} + n_{Ar}$ is the total moles of oxygen plus argon in the gas and P is the total pressure. Combining (3) and (5), one gets

$$C_{O_2} = \frac{dn_{O_2}}{d\mu_{O_2}} = \frac{p_{O_2}PV}{(RT)^2 p_{Ar}} = \frac{p_{O_2}PV}{(RT)^2 (P - p_{O_2})} = \frac{p_{O_2}V}{(RT)^2 [1 - (p_{O_2}/P)]} \quad (6)$$

When $p_{O_2} \ll P$, the thermodynamic capacity of oxygen is proportional to the oxygen partial pressure. As p_{O_2} tends to P , the infinite capacity of pure oxygen is approached.

In the numerical procedure for the calculation of C_{O_2} , the relative chemical potential is computed for an initial mixture of n_{O_2} moles of O_2 and n_{Ar} moles of Ar using (4). The chemical potential is then changed by 100 or 1000 cal by addition or removal of oxygen. This changes the number of moles of oxygen in the system by Δn_{O_2} , which is given by

$$\Delta n_{O_2} = \frac{[p'_{O_2}(n_{Ar} + n_{O_2}) - n_{O_2}]}{(1 - p'_{O_2})}, \quad (7)$$

where p'_{O_2} is the new oxygen partial pressure, corresponding to μ'_{O_2} . The thermodynamic capacity is given by

$$C_{O_2} = \Delta n_{O_2} / (\mu'_{O_2} - \mu_{O_2}). \quad (8)$$

The calculations have been carried out per mole of the mixture. The variation of $\log C_{O_2}$ with $\log p_{O_2}$ is shown in figure 1 at 1 atm total pressure for different temperatures. Both analytical and numerical techniques give essentially the same result.

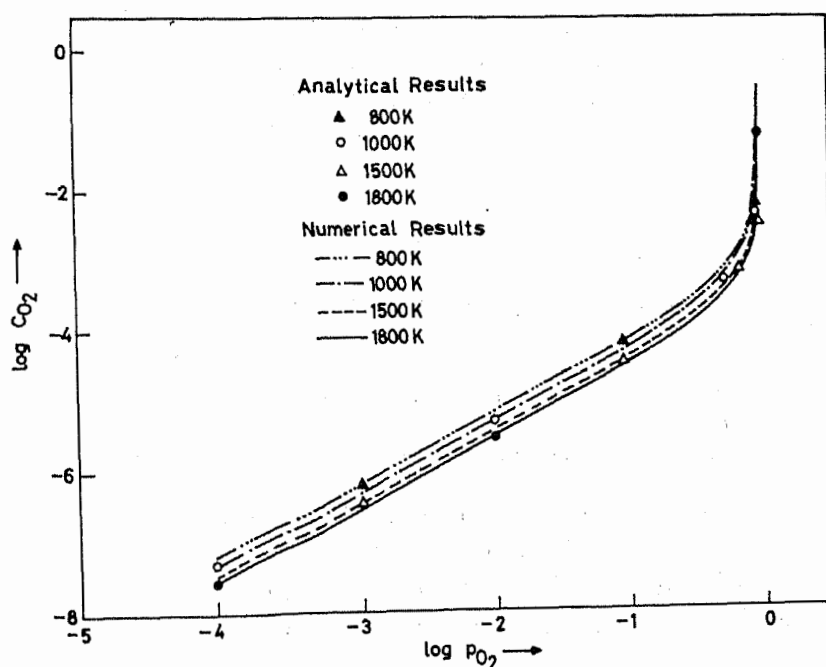


Figure 1. Variation of the thermodynamic capacity, C_{O_2} , with p_{O_2} in $Ar + O_2$ gas mixture at different temperatures.

4. Reactive gas mixtures containing stable species

4.1 CO + CO₂ mixtures

In a CO + CO₂ gas mixture, the system is buffered against changes in oxygen potential by the formation or dissociation of CO₂ according to the reaction



The equilibrium constant K is given by

$$K = (p_{\text{CO}}/p_{\text{CO}_2})p_{\text{O}_2}^{1/2}. \quad (10)$$

Based on the approach similar to the one used for an Ar + O₂ mixture, the thermodynamic capacity of oxygen, C_{O_2} , at a given oxygen partial pressure p_{O_2} can be approximated by

$$C_{\text{O}_2} = \frac{V}{(RT)^2} \left[p_{\text{O}_2} + \frac{P(p_{\text{O}_2}/K^2)^{1/2}}{4[1+(p_{\text{O}_2}/K^2)^{1/2}]^2} \right] \frac{1}{(1-p_{\text{O}_2}/P)}. \quad (11)$$

At $p_{\text{O}_2} \simeq K^2$ or a $p_{\text{CO}}/p_{\text{CO}_2}$ ratio equal to unity, C_{O_2} is a maximum. A minimum in capacity is observed at a value of p_{O_2} given by

$$p_{\text{O}_2} \simeq \frac{1}{4}(P^{2/3} K^{2/3}), \quad (12)$$

or

$$\frac{p_{\text{CO}_2}}{p_{\text{CO}}} \simeq \frac{1}{4}(P^{1/3} K^{-2/3}). \quad (13)$$

In the numerical technique, we consider a starting gaseous mixture of carbon dioxide, carbon monoxide and oxygen with initial number of moles of the species being n_{CO_2} , n_{CO} and n_{O_2} respectively. The equilibrium composition at a given temperature may be reached by the decomposition of a certain amount of CO₂, say α , according to the reaction (9). Each of the partial pressures in (10) may be expressed in terms of mole fractions of the particular species. For example

$$p_{\text{CO}_2} = \frac{n_{\text{CO}_2} - \alpha}{n_{\text{CO}_2} + n_{\text{CO}} + n_{\text{O}_2} + \alpha/2}. \quad (14)$$

Substituting for the partial pressures in (10), the value of α is given by the solution of the equation

$$\alpha^3(1-K^2)/2 + \alpha^2(1-K^2)(n_{\text{O}_2} + n_{\text{CO}}) + \alpha[K^2 n_{\text{CO}_2}(3/2 n_{\text{CO}_2} + 2n_{\text{CO}} + 2n_{\text{O}_2}) + n_{\text{CO}}(n_{\text{CO}}/2 + 2n_{\text{O}_2})] - [K^2 n_{\text{CO}_2}^2(n_{\text{CO}_2} + n_{\text{CO}} + n_{\text{O}_2}) - n_{\text{CO}}^2 n_{\text{O}_2}] = 0. \quad (15)$$

Let Δn_{O_2} be the change in total number of moles of oxygen corresponding to a change in chemical potential by 1000 calories. Equilibrium is reached by the formation of x additional moles of CO₂. The new values of the partial pressures are represented as p'_{CO_2} , p'_{CO} and p'_{O_2} . The values of x and Δn_{O_2} are then given by the expressions

$$x = (n_{\text{CO}} + \alpha) - \frac{p'_{\text{CO}}(n_{\text{CO}_2} + n_{\text{CO}})}{(1-p'_{\text{O}_2})} \quad (16)$$

and

$$\Delta n'_{O_2} = \left(\frac{n_{CO}}{2} - n_{O_2} \right) - \frac{(n_{CO_2} + n_{CO})(p'_{CO} - 2p'_{O_2})}{2(1 - p'_{O_2})} \quad (17)$$

Hence

$$C_{O_2} = \Delta n'_{O_2} / \Delta \mu_{O_2} \quad (18)$$

Two separate cases of $CO + CO_2 + O_2$ mixtures can be considered. In the first case the equilibrium O_2 concentration is generated solely by the dissociation of CO_2 in a $CO + CO_2$ mixture. For this case, initial oxygen $n_{O_2} = 0$. In the second case, the equilibrium CO is generated by dissociation of CO_2 in a $CO_2 + O_2$ mixture and therefore $n_{CO} = 0$.

Case 1: When $n_{O_2} = 0$, the expressions (15) and (17) reduce to

$$\begin{aligned} & \alpha^3(1 - K^2)/2 + \alpha^2 n_{CO}(1 - K^2) + \alpha[(3/2)K^2 n_{CO_2}^2 + 2K^2 n_{CO_2} n_{CO} + n_{CO}^2/2] \\ & - K^2 n_{CO_2}^2 (n_{CO_2} + n_{CO}) = 0, \end{aligned} \quad (19)$$

$$\Delta n_{O_2} = \frac{n_{CO}}{2} - \frac{(n_{CO_2} + n_{CO})(p'_{CO} - 2p'_{O_2})}{2(1 - p'_{O_2})} \quad (20)$$

When n_{CO} is large, p'_{CO} is also large, therefore $\Delta n'_{O_2}$ is small. As n_{CO} decreases, p'_{CO} also decreases. However, $\Delta n'_{O_2}$ increases with decreasing n_{CO} . As the ratio n_{CO_2}/n_{CO} approaches unity, the value of $\Delta n'_{O_2}$ and hence the thermodynamic capacity increases. A maximum value is obtained at a ratio of one. With increasing value of the n_{CO_2}/n_{CO} ratio, capacity decreases and passes through a minimum. It is difficult to control oxygen potential near the minimum in capacity. Other systems should be considered if it is necessary to control oxygen potential in this range.

Case 2: When $n_{CO} = 0$, i.e. at very high oxygen potentials, decomposition of CO_2 alone cannot account for the free oxygen in the system. The starting mixture would therefore have CO_2 and O_2 as the initial species, the CO content at equilibrium being set up by dissociation of CO_2 . Then α and $\Delta n'_{O_2}$ are given by the solutions of

$$\begin{aligned} & \alpha^3(1 - K^2)/2 + \alpha^2(1 - K^2)n_{O_2} + \alpha[(3/2)K^2 n_{CO_2}^2 + 2K^2 n_{CO_2} n_{O_2}] \\ & - K^2 n_{CO_2}^2 (n_{CO_2} + n_{O_2}) = 0 \end{aligned} \quad (21)$$

and

$$\Delta n'_{O_2} = \frac{(2p'_{O_2} - p'_{CO})n_{CO_2}}{2(1 - p'_{O_2})} - n_{O_2} \quad (22)$$

respectively. As n_{O_2} and p'_{O_2} increase, the thermodynamic capacity of oxygen increases. For pure oxygen, the capacity will be infinite.

Figure 2 gives the thermodynamic capacity of oxygen as a function of the equilibrium n_{CO_2}/n_{CO} ratio per mole of gas mixture. The equilibrium constants were taken from the Janaf tables (Stull and Prophet 1971). The values given by the analytic expression (11) and the numerical technique match well. The plot shows a peak at a 1:1 ratio, where the system is buffered well against changes in oxygen potential. At higher oxygen potentials, the capacity coincides with that of an $Ar + O_2$ mixture as shown in figure 3. The capacity values at 1500 K are tabulated for various input compositions in table 1.

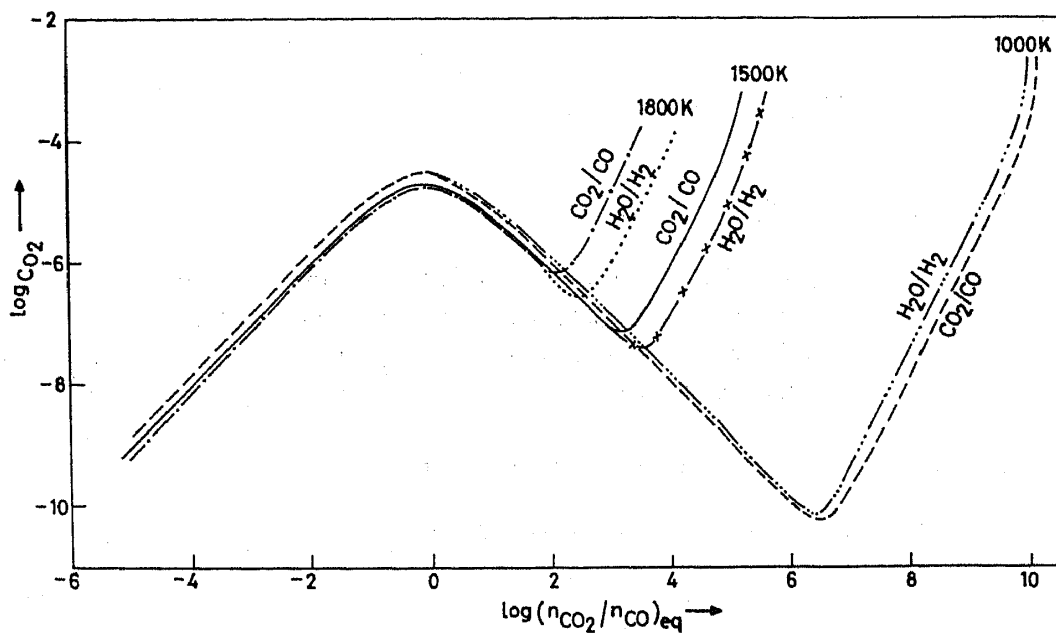


Figure 2. Variation of C_{O_2} with the equilibrium n_{CO_2}/n_{CO} or n_{H_2O}/n_{H_2} ratios at various temperatures.

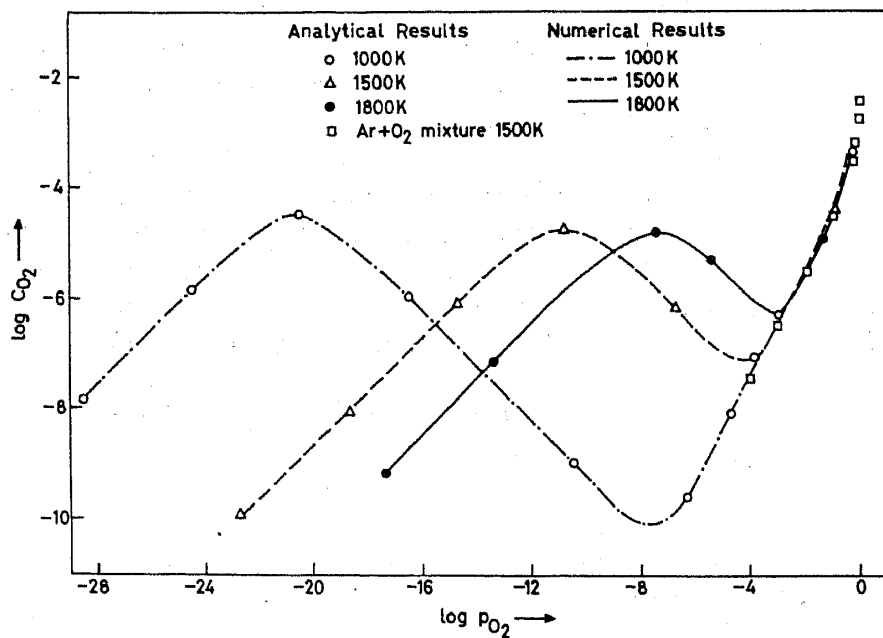


Figure 3. Variation of C_{O_2} as a function of $\log p_{O_2}$ in a $CO + CO_2$ gas mixture at various temperatures.

4.2 $H_2 + H_2O$ mixtures

The procedure for calculating C_{O_2} is similar to that for a $CO + CO_2$ gas mixture. As seen in figure 2, the small differences in the values of the equilibrium constants for the dissociation reactions in the two gas mixtures give rise to significant differences in the values of C_{O_2} only at higher oxygen potentials.

Table 1. Values of thermodynamic capacity of oxygen in a CO + CO₂ mixture for defined values of the equilibrium ratio $n_{\text{CO}_2}/n_{\text{CO}}$ in the input mixture at 1500 K, obtained by the numerical technique.

$n_{\text{CO}_2}/n_{\text{CO}}$	p_{O_2} atm	C_{O_2} moles/cal
10^{-7}	2.13×10^{-25}	1.19×10^{-10}
10^{-5}	2.13×10^{-21}	8.94×10^{-10}
10^{-3}	2.13×10^{-17}	9.15×10^{-8}
10^{-1}	2.13×10^{-13}	7.45×10^{-6}
1	2.13×10^{-11}	2.10×10^{-5}
10	2.13×10^{-9}	6.49×10^{-6}
2.59×10^3	1.43×10^{-4}	8.72×10^{-8}
2.16×10^4	9.92×10^{-3}	4.03×10^{-6}
6.53×10^4	9.09×10^{-2}	4.17×10^{-5}
1.53×10^5	0.5	6.68×10^{-4}

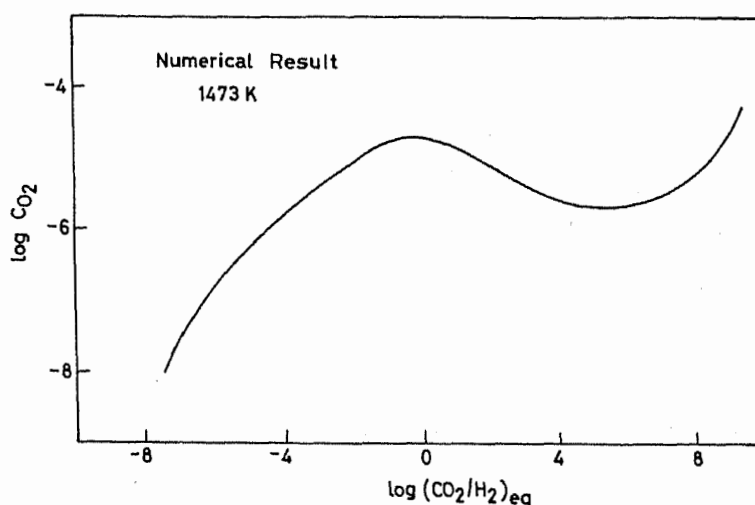


Figure 4. Dependence of C_{O_2} on $n_{\text{CO}_2}/n_{\text{H}_2}$ at 1473 K.

4.3 $\text{H}_2 + \text{CO}_2$ mixtures

The equilibrium gas composition at 1473 K was calculated considering the species CO₂, CO, H₂O, H₂ and CH₄, using the SOLGAS program (Eriksson 1971). The oxygen capacity was then computed using the numerical procedure. As seen from figure 4, a peak is seen at a ratio of $n_{\text{CO}_2}/n_{\text{H}_2}$ equal to unity, at which the equilibrium $n_{\text{CO}_2}/n_{\text{CO}}$ and $n_{\text{H}_2\text{O}}/n_{\text{H}_2}$ values are also close to one. The trend is similar to that for other reactive gas mixtures.

5. Reactive gas mixtures containing unstable species

5.1 $\text{SO}_2 + \text{SO}_3$ mixtures

The equilibrium constant for the decomposition of SO₃ is large even at a low temperature of 1000 K. Figure 5 gives the C_{O_2} values in an SO₂ + SO₃ mixture as a

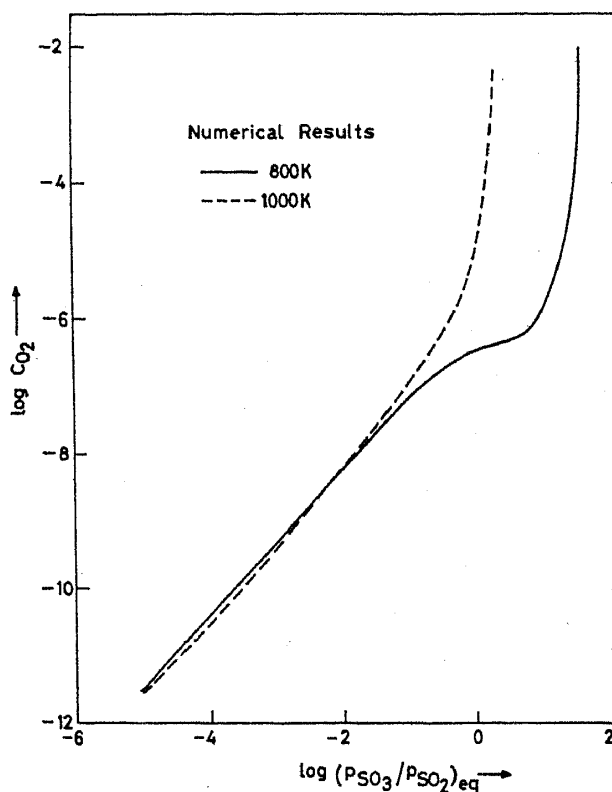


Figure 5. Variation of C_{O_2} with equilibrium n_{SO_3}/n_{SO_2} in an $SO_2 + SO_3$ gas mixture at 800 K and 1000 K.

function of the equilibrium n_{SO_3}/n_{SO_2} ratio. No peak is observed even at 1000 K and it is not clearly defined at 800 K. For identical ratios of n_{SO_3}/n_{SO_2} and n_{CO_2}/n_{CO} at a given temperature, the thermodynamic capacity of oxygen in an $SO_2 + SO_3$ mixture is lower.

5.2 $H_2 + H_2S$ mixtures

The thermodynamic capacity of sulphur C_{S_2} in a $H_2 + H_2S$ gas mixture at temperatures of 800, 1000, 1500 and 1800 K has been calculated on similar lines. The peak again occurs at a ratio of unity as seen in figure 6. At higher temperatures where H_2S is unstable, no peak is seen and the capacity continuously increases.

6. Condensed phase calculations

In the copper-oxygen system, when oxygen is added to molten copper, its chemical potential rises. However, more oxygen is required at higher potentials to produce a given rise and the capacity increases with increasing oxygen content. While the mathematical equation referring to this curve has no upper limit, there are very real chemical limits. When the oxygen potential exceeds $-32,860$ cal at 1373 K, Cu_2O will precipitate as a separate phase. Attempts to increase the oxygen content of molten copper will result in the formation of Cu_2O , which is equivalent to a "leak" at

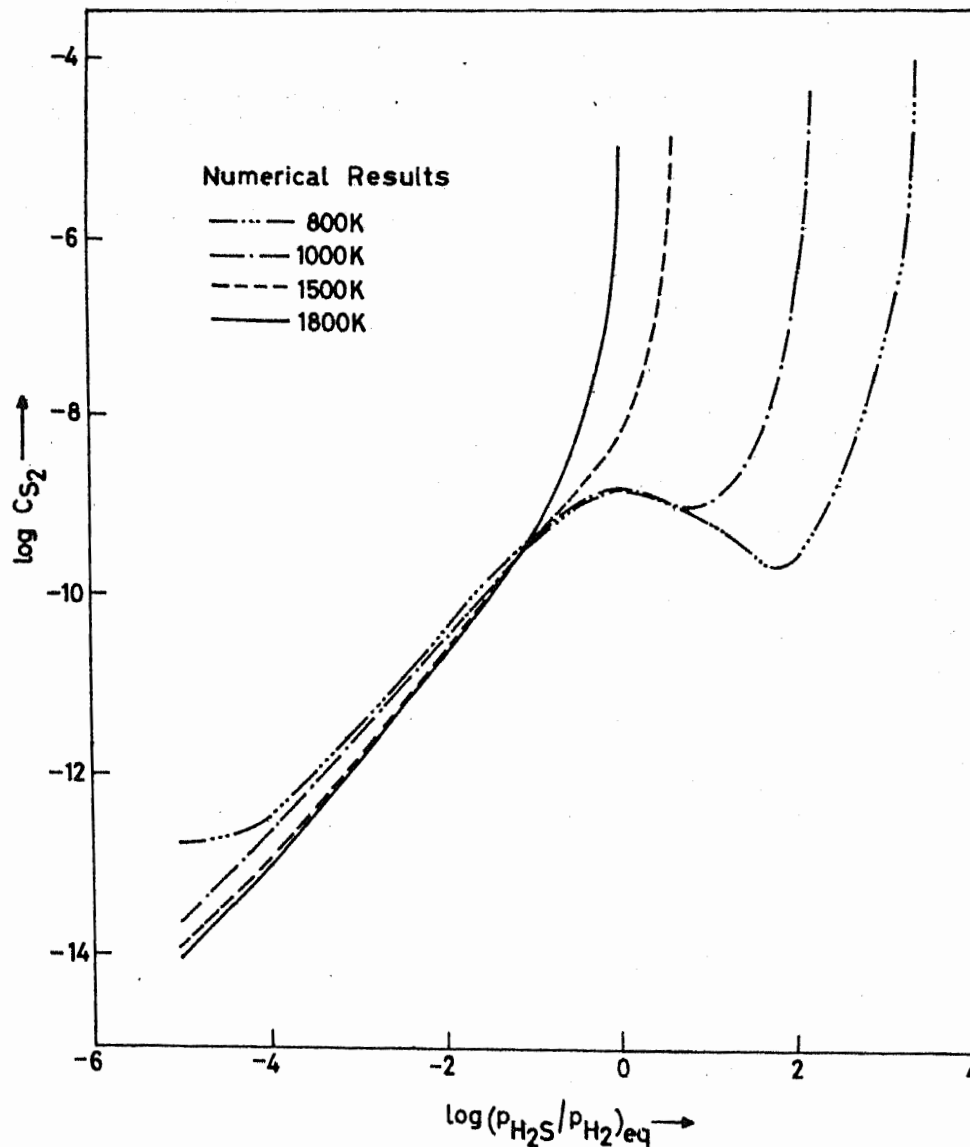


Figure 6. Thermodynamic capacity of sulphur in an $H_2 + H_2S$ gas mixture at different temperatures.

this level into another phase with unlimited capacity. The whole copper can be oxidized at this potential.

Similarly, if copper contained dissolved tin, it would not be possible to raise the oxygen potential above the value corresponding to the formation of SnO_2 . The oxygen potential for the formation of SnO_2 will be a function of the tin concentration. Here, the leak is to another oxide SnO_2 . As the tin concentration decreases due to the formation of SnO_2 the permissible oxygen potential in the metal goes up. Since the oxygen potential does not remain constant during the formation of SnO_2 , the capacity does not go to infinity as in the $Cu-O$ system, but increases to a very large value at the saturation solubility limit for oxygen at the given tin concentration.

6.1 Molten Cu-O system

Dissolution of oxygen in molten copper may be represented by the reaction



where ΔG° , the free energy change associated with the above reaction is given by (Jacob and Jeffes 1971)

$$\Delta G^\circ = RT \left(\frac{\ln(\% O f_O^\circ)}{p_{O_2}^{1/2}} \right) = -20,542 + 1.723 T \text{ cal/g.at.} \quad (24)$$

The standard state for dissolved oxygen is chosen such that $a_O = \text{atom \% O as atom \% O} \rightarrow 0$. The variation of the Henrian activity coefficient of oxygen with concentration at 1373 K is given by

$$\log f_O^\circ = -0.042 \text{ (at \% O)} \quad (25)$$

Differentiating (24) with respect to the change in number of moles of oxygen, dn_{O_2} , the thermodynamic capacity, C_{O_2} , is given by the expression

$$C_{O_2} = \frac{dn_{O_2}}{d\mu_{O_2}} = \frac{X_{O_2}}{2RT(1 + 2\varepsilon_O^\circ X_{O_2})}, \quad (26)$$

where X_{O_2} is the oxygen content in molten copper expressed as a mole fraction of O_2 and ε_O° is the self interaction parameter, $(\partial \ln f_O / \partial X_O)_{X_O \rightarrow 0}$.

In the numerical technique, the equilibrium constant K for (23) given by

$$K = a_O / p_{O_2}^{1/2} \quad (27)$$

was used to calculate the oxygen partial pressure in equilibrium with a given amount of dissolved oxygen in molten copper. The chemical potential of oxygen was then incremented by 100 cal and the new oxygen content computed. The analytic and numerical results agree well. The saturation solubility limit for oxygen corresponding to the formation of Cu_2O at 1373 K is 2.4 atom % O. As seen from figure 7, the

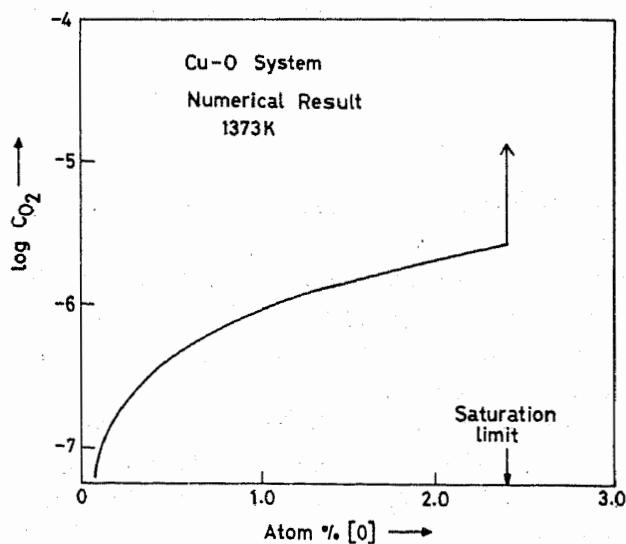


Figure 7. C_{O_2} as a function of dissolved oxygen in molten copper at 1373 K.

thermodynamic capacity, C_{O_2} , plotted as a function of dissolved oxygen in copper expressed in atom % becomes infinite at the saturation value. The phase mixture Cu-O + Cu₂O provides a composition invariant buffer for oxygen.

6.2 Molten Cu-O-Sn system

The calculations were done at 1373 K for various alloy compositions in the range 1-14 atom % Sn. The interaction parameter ϵ_O^{Sn} is (Jacob *et al* 1970)

$$\epsilon_O^{Sn} = \left(\frac{\partial \ln f_O}{\partial X_{Sn}} \right)_{X_{Sn}, X_O \rightarrow 0} = -4.6. \quad (28)$$

The corresponding expression for $\log f_O$ is

$$\log f_O = -0.042 (\text{atom \% O}) - \frac{4.6}{2.303} X_{Sn}. \quad (29)$$

The activity coefficient of tin in molten copper is given by (Jacob *et al* 1970)

$$\log \gamma_{Sn} = (-1.38 + 3.333 X_{Sn}) X_{Cu}^2 \quad (30)$$

Combining (24) with the free energy of formation of SnO₂, the free energy change for the reaction



is

$$\Delta G^\circ = -98,566 + 47.174 T \text{ cal/mol} \quad (32)$$

The upper limit to the oxygen activity corresponding to saturation can thus be calculated. The capacity values for various oxygen contents upto the saturation limit were computed by the numerical technique. At saturation, the formation of SnO₂ provides an additional sink for oxygen. The analytic expression for capacity of oxygen in saturated solutions of Cu-Sn is obtained as shown below:

$$\begin{aligned} \frac{d\mu_{O_2}}{dX_{Sn}} &= \frac{d}{dX_{Sn}} (\ln X_{Sn} \gamma_{Sn}) \\ &= -\frac{RT}{X_{Sn}} - RT \left(\frac{d \ln \gamma_{Sn}}{dX_{Sn}} \right). \end{aligned} \quad (33)$$

$$\text{Now } dX_{Sn} = -dX_{SnO_2} = -dn_{O_2}. \quad (34)$$

Hence

$$C_{O_2} = \frac{dn_{O_2}}{d\mu_{O_2}} = \frac{X_{Sn}}{RT\{1 + 2.303 X_{Sn}(1 - X_{Sn})(2.76 + 3.333[1 - 3X_{Sn}])\}}. \quad (35)$$

At saturation the rise in capacity is by 2-3 orders of magnitude as shown in figure 8. The capacity of the two-phase mixture of SnO₂ and the homogeneous solution is larger than that of the single phase alloy. The saturation oxygen concentration in the ternary alloy increases as the tin concentration decreases. The oxygen capacity therefore does not go to infinity at saturation. As the tin concentration in the alloy is

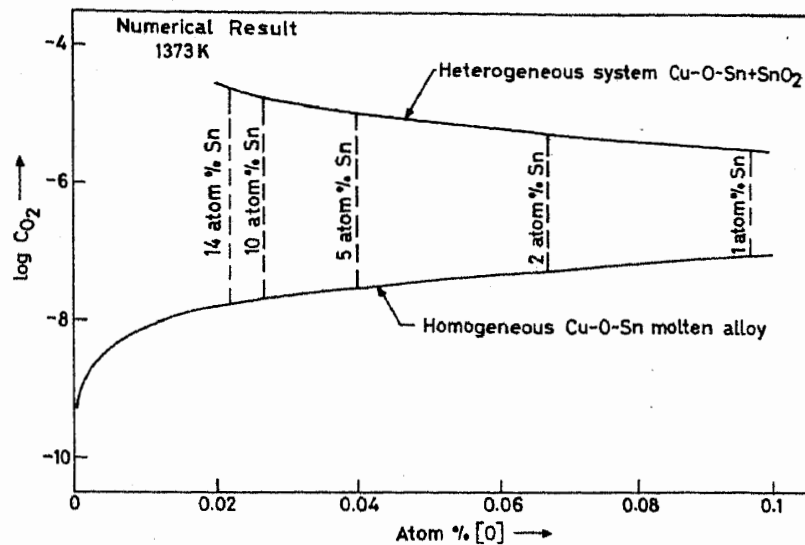


Figure 8. Variation of C_{O_2} with the concentration of dissolved oxygen in Cu-Sn alloys of different compositions at 1373 K.

depleted by the formation of SnO_2 , the thermodynamic capacity of oxygen varies along the upper curve corresponding to the two-phase mixture. This example illustrates a composition variant two phase buffer.

7. Applications

In the preparation of inorganic solids of controlled compositions, the samples are often equilibrated with a gas phase of controlled chemical potential. If the thermodynamic capacity of the gas is low, then a large volume of the gas is required unless the amount of oxygen to be exchanged is small. Similarly, when using gas mixtures containing molecules that differ considerably in molecular weights, as in $H_2 + H_2S$ mixtures, the flow rate must be kept high to minimize thermal segregation. The experiments using large volumes of high purity gases are inherently expensive. Sometimes disposal of large volumes of toxic gases can also be problematic. A solution to the problem is to use a recirculating system which provides for regeneration of the gas to the required composition. Solid state sensors can be placed in the recirculating stream to ensure that the composition is maintained at preset values. In gas mixtures, for the control of oxygen potential, injection or withdrawal of oxygen from the stream can be accomplished by an electrochemical oxygen pump. Coulometric oxygen pumps depend on Faraday's law relating current to oxygen flux and are based on solid electrolytes with mobile oxygen ions. In principle the pump is similar to a sensor. The sensor operates in open circuit mode while the pump operates in the non-zero current mode. The current through the oxygen pump can be regulated by the emf feedback from the solid state sensor. The feedback amplification may be automated. It can be shown that the correct value of the pump current is related to the thermodynamic capacity of oxygen in the gas

$$I = \frac{4Fn_V}{1-t_e} \int_{\mu_{O_2}'}^{\mu_{O_2}''} \frac{C_{O_2}}{V} d\mu_{O_2} \quad (36)$$

where n_V is the volume flow rate of the gas, μ'_{O_2} and μ''_{O_2} are the actual and preset chemical potentials of oxygen in the gas, \bar{t}_e is the average electronic transference number of the solid electrolyte and C_{O_2}/V is the oxygen capacity of the gas per unit volume. The pumping current is a highly nonlinear function of oxygen partial pressure.

8. Conclusion

A new thermodynamic variable, thermodynamic capacity, which is a measure of the ability of a chemical system to exchange a species without a change in its chemical potential, has been proposed. The variation of thermodynamic capacity for ideal and non-ideal solutions and systems with strong and weak compounds have been computed using analytical and numerical techniques. The capacities in two-phase heterogeneous systems are illustrated using $Cu-O + Cu_2O$ and $Cu-O-Sn + SnO_2$ as examples. Multiphase systems can be devised that have higher thermodynamic capacity than homogeneous phases. The concept of thermodynamic capacity is useful in high temperature chemical vapour transport and solid state devices incorporating electrochemical pumps.

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